Inside Cover	Costa's Levels
INB Score Sheet	of THINKING
	Level 1 - Who, What, Where, Why, How
	Level 2 - Explain Compare Contrast
	Level 3 - Predict Apply Evaluate

Left Side	Right Side
OUTPUT	INPUT
-Summaries	-Notes
- -pagers -Practice Problems	- Whiteboard Pictures -Lab Notes

Table	of Contents	Table of Cont	ents	

Table	of Contents	Table of Cont	ents	

SWBAT: Model taking Cornell Notes	
1	

	SWBAT define Chemistry and explain how to use a mass balance
	2 L1 Q
	l L3 Q
2	Sum: 3-5 sent.

	Mass and Change Lab Summary		
Ч		5	
	•		

	SWBAT explain Law of Conservation of Mass and Physical vs. Chemical Chgs.	
	3 Qs 2L2	
6	Svmmary 7	

	SWBAT compare and Contrast between Accuracy and Precision.	
	3 L2 Qs	
	Summany	
8	9	

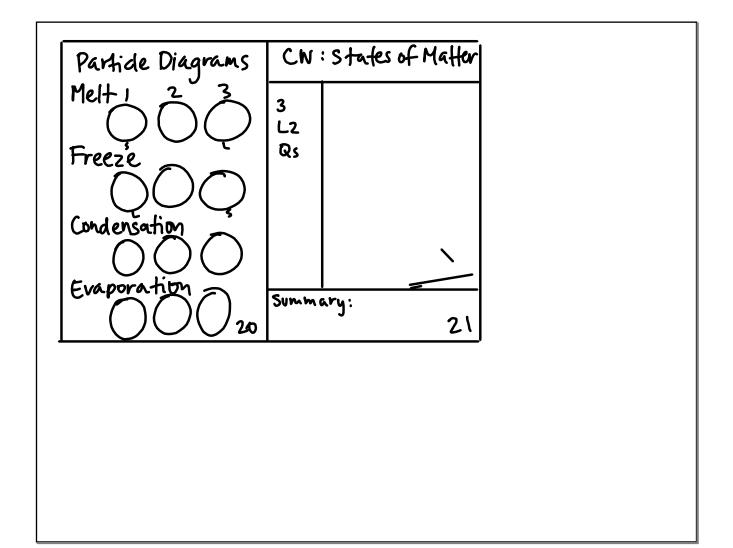
Notes	s: Metric Prefixes	SWi Vol	BAT define mass, lume and Density.
ВІЧ	Kilo-1000 Base-gram Liter meter	3 L2 Qs	•
Small	Centi - 1/1000 Milli - 1/1000		
10		Summ	nanz.

	u2 KW	L	
	What Do Wi You KNOW? V to	hat Do What DI Did You Know? LEARN!	
12		13	

Demo	Observations	Whitele	poarding
PopCorn Demo		١	2
Demo		3	T
Svmma	ny: See website for Instructions	Note:	See Webgite for Notes
14			(5

Demo	Observations	Whiteboarding		
Water Dye		1	2	
Dye Demo		3	1	
Svmma	ry: See website for Instructions	Note:	See Webgite for Notes	
16				17

-pager of Demos • One-page Summany of Notes and what you learned from	Video Notes: Molecules in Solids	
fne 2 demos Must Include: -Color	Molecules in Liquids	
-Border -3 pictures	Evaporation of Condensation	4
-2 Level 2 Qs with Answers	19_	



Demo! Expansion of Liquids	Expansion and Contraction
Observations:	
Video & Demo Summany:	Heat & Temperature
22	23

Practice Problems P,V #	P,V,T,n Notes	
P,n 🗡	See next SLIDE for Set-up	
P,T		
24	25	

Parameter	and Units	Relationship & Explanation	Picture
Pressure (P)	Volume =amount of space	ΛV , $\Delta P = inverse$ ΔV , $\Lambda P = ifless space$	= latin
is the	Particles	then particles will collide more often. c=n,T 1 n, TP = direct	= atm
particles have with	=puffs mL= mL=	when you increase the # of air particles in a constant volume they will collide more T pressure C = V, T	= 2atm
container. Units atm=1	Temp	,	
mm Hg=760 KPg=101.3			25

A 2.0L container at 25°C has 25 puffs of air at a pressure of 800 mm Hg. The container is opened and the pressure drops to 710 mm Hg. How many puffs escaped?

P	n	Show work
Initial 800 to 25 puffs	Show work	
Final 10 to 25 puffs	Final 10 to 25 puffs	South 10 = 25 puffs x 700 = 22.2 puffs
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- 1) Explain Problem (Introduce)
- 2) Explain Table
- 3) Explain Calculation 4) Does this answer make sense?